

Q#	Equation	Shift Left or Right?	Changes?
1	$\text{N}_{2(g)} + \text{O}_{2(g)} \leftrightarrow 2\text{NO}_{(g)}$		
	Stressor:		
2	$\text{H}_{2(g)} + \text{I}_{2(g)} \leftrightarrow 2\text{HI}_{(g)}$		
	Stressor:		
3	$\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \leftrightarrow \text{CO}_{2(g)} + \text{H}_{2(g)}$		
	Stressor:		
4	$2\text{SO}_{2(g)} + \text{O}_{2(g)} \leftrightarrow 2\text{SO}_{3(g)}$		
	Stressor:		
5	$3\text{O}_{2(g)} \leftrightarrow 2\text{O}_{3(g)}$		
	Stressor:		
6	$\text{H}_2\text{O}_{2(l)} \leftrightarrow \text{H}_{2(g)} + \text{O}_{2(g)}$		
	Stressor:		
7	$\text{CO}_{(g)} + 2\text{H}_{2(g)} \leftrightarrow \text{CH}_3\text{OH}_{(g)}$		
	Stressor:		
8	$\text{CH}_{4(g)} + 2\text{O}_{2(g)} \leftrightarrow \text{CO}_{2(g)} + 2\text{H}_2\text{O}_{(g)}$ $\Delta H = -5\text{kJ}$		
	Stressor:		